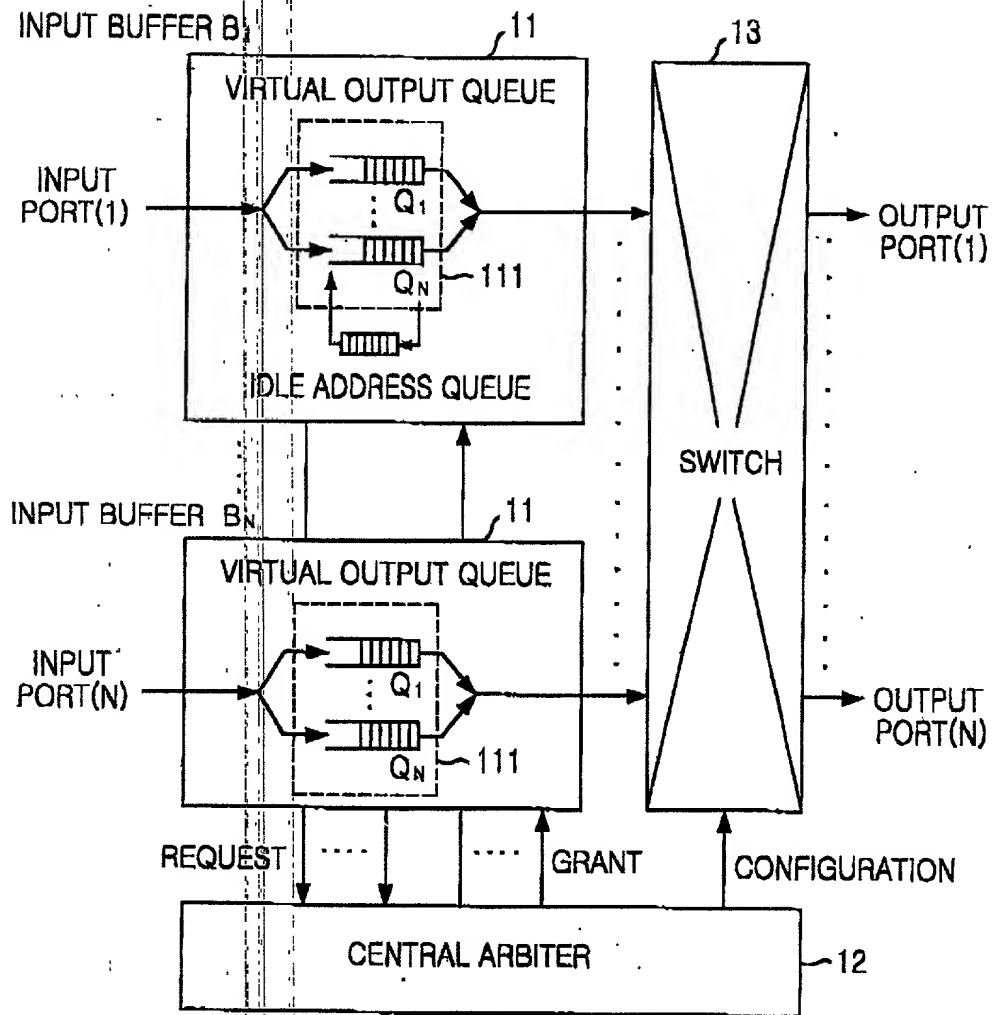


FIG. 1



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FIG. 2

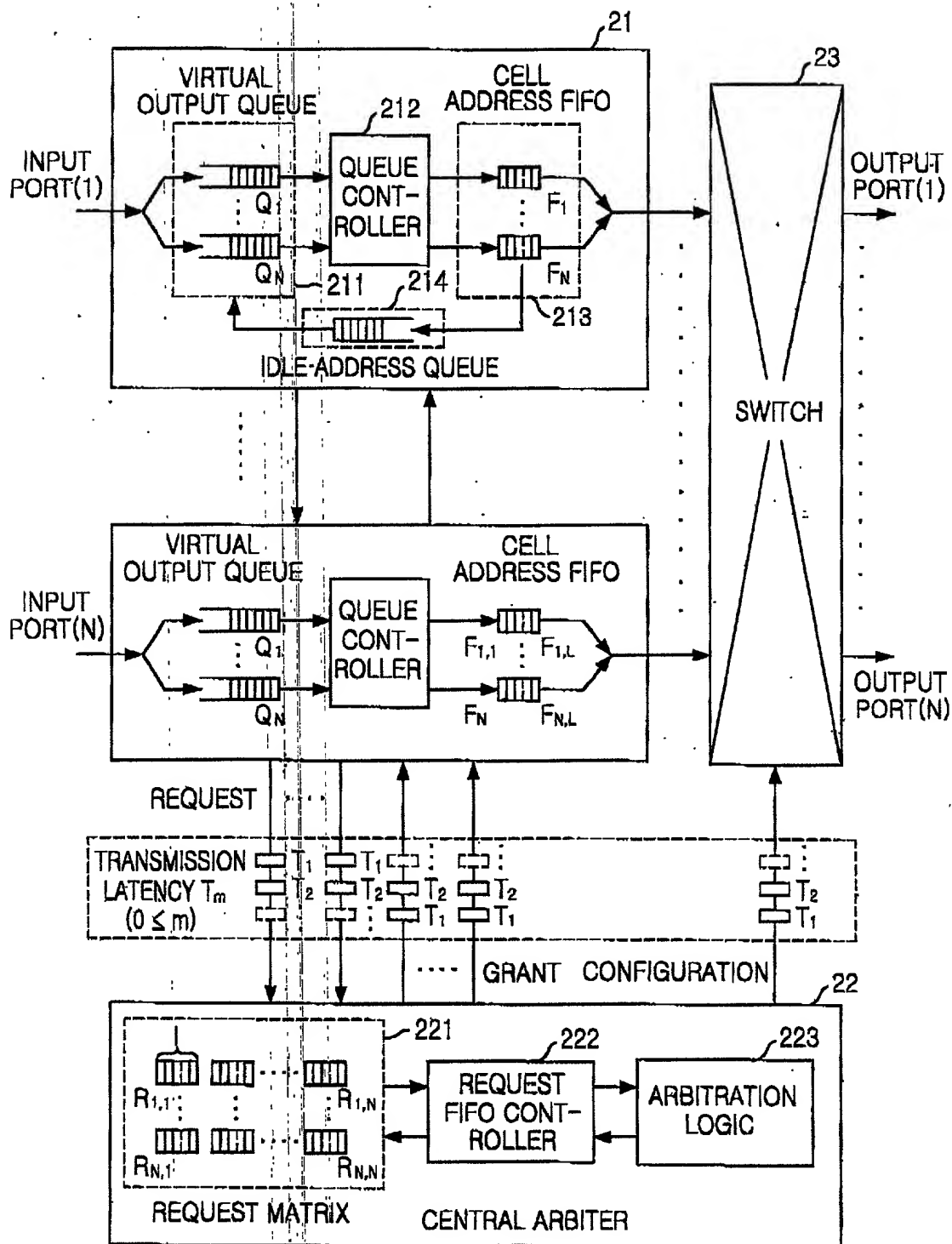


FIG. 3

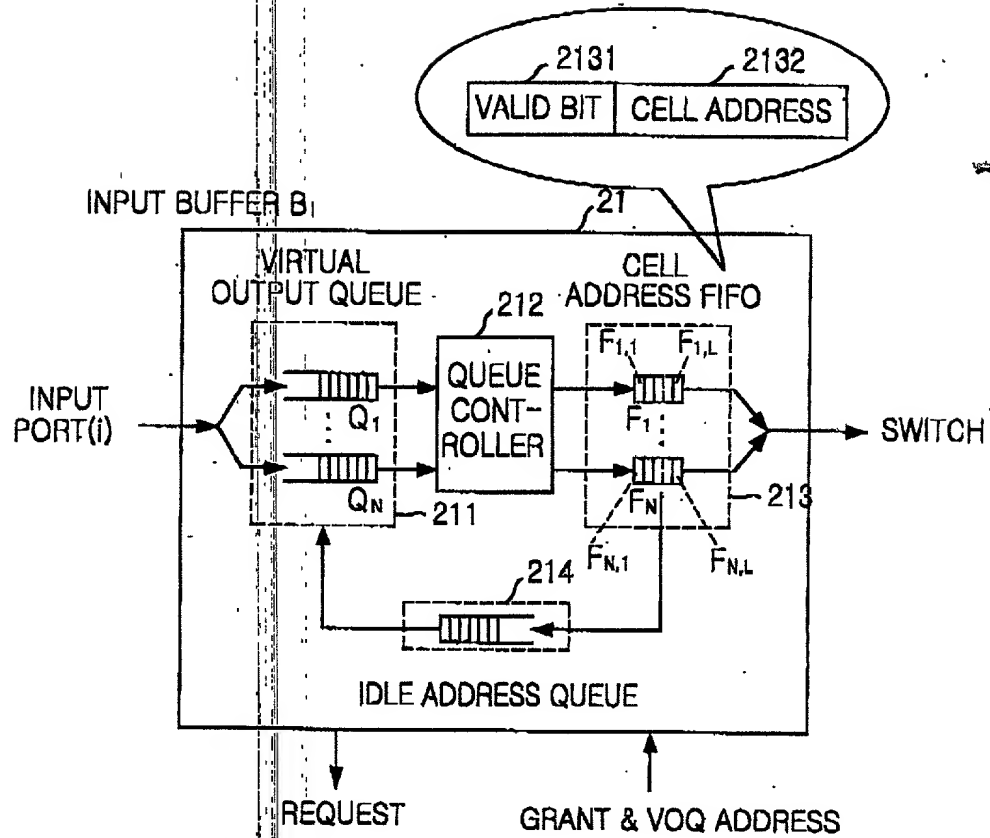
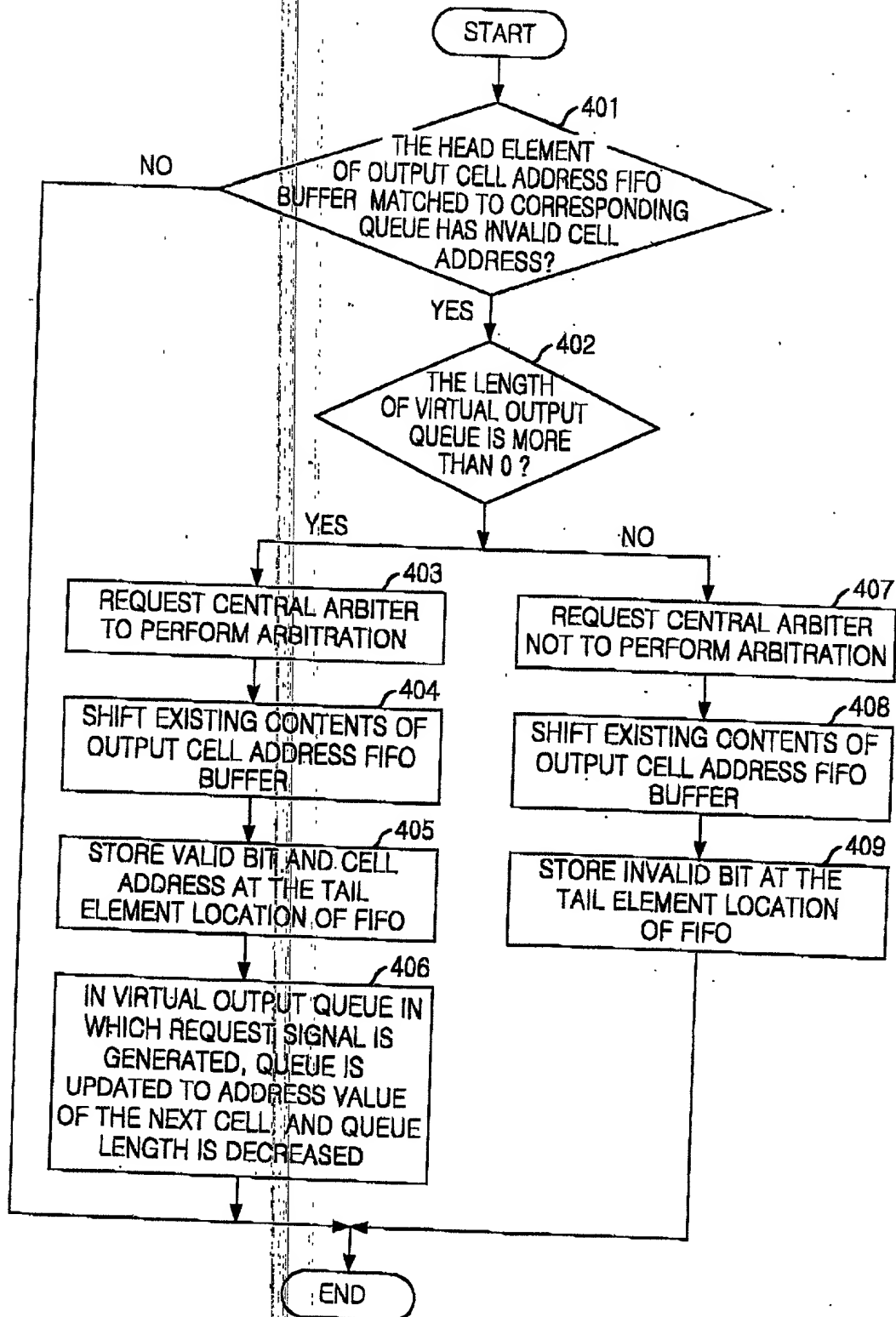


FIG. 4



Parameter	Value	Unit
Initial concentration of H_2O_2	0.01	M
Initial concentration of Fe^{2+}	0.001	M
Initial concentration of H^+	0.1	M
Temperature	25	$^\circ\text{C}$
Reaction time	0-100	min
Reaction rate	0.001	$\text{M}^{-1}\text{s}^{-1}$
Reaction order	1	
Reaction mechanism	Free radical chain reaction	
Reaction products	Fe^{3+} , H_2O , O_2	
Reaction conditions	Dark, sealed, stirred	
Reaction medium	Aqueous solution	
Reaction vessel	100 mL glass bottle	
Reaction setup	Stirrer, thermometer, gas outlet	
Reaction analysis	UV-Vis, titration, gas analysis	
Reaction results	Complete reaction within 100 min	
Reaction conclusion	Reaction is first order with respect to H_2O_2	
Reaction equation	$\text{H}_2\text{O}_2 + \text{Fe}^{2+} + \text{H}^+ \rightarrow \text{Fe}^{3+} + \text{H}_2\text{O} + \text{O}_2$	
Reaction rate equation	$\text{Rate} = k[\text{H}_2\text{O}_2]$	
Reaction rate constant	0.001	min^{-1}
Reaction rate half-life	69.3	min
Reaction rate activation energy	50.0	kJ/mol
Reaction rate pre-exponential factor	1.0	min^{-1}
Reaction rate Arrhenius equation	$\ln k = \ln A - E_a/RT$	
Reaction rate Arrhenius plot	Linear plot of $\ln k$ vs $1/T$	
Reaction rate Arrhenius slope	-5000	K
Reaction rate Arrhenius intercept	10.0	$\ln \text{min}^{-1}$
Reaction rate Arrhenius correlation coefficient	0.99	
Reaction rate Arrhenius equation	$k = 1.0 \times 10^4 \exp(-5000/T)$	
Reaction rate Arrhenius plot	Linear plot of $\ln k$ vs $1/T$	
Reaction rate Arrhenius slope	-5000	K
Reaction rate Arrhenius intercept	10.0	$\ln \text{min}^{-1}$
Reaction rate Arrhenius correlation coefficient	0.99	
Reaction rate Arrhenius equation	$k = 1.0 \times 10^4 \exp(-5000/T)$	
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Reaction rate Arrhenius correlation coefficient	0.99	
Reaction rate Arrhenius equation	$k = 1.0 \times 10^4 \exp(-5000/T)$	
Reaction rate Arrhenius plot		

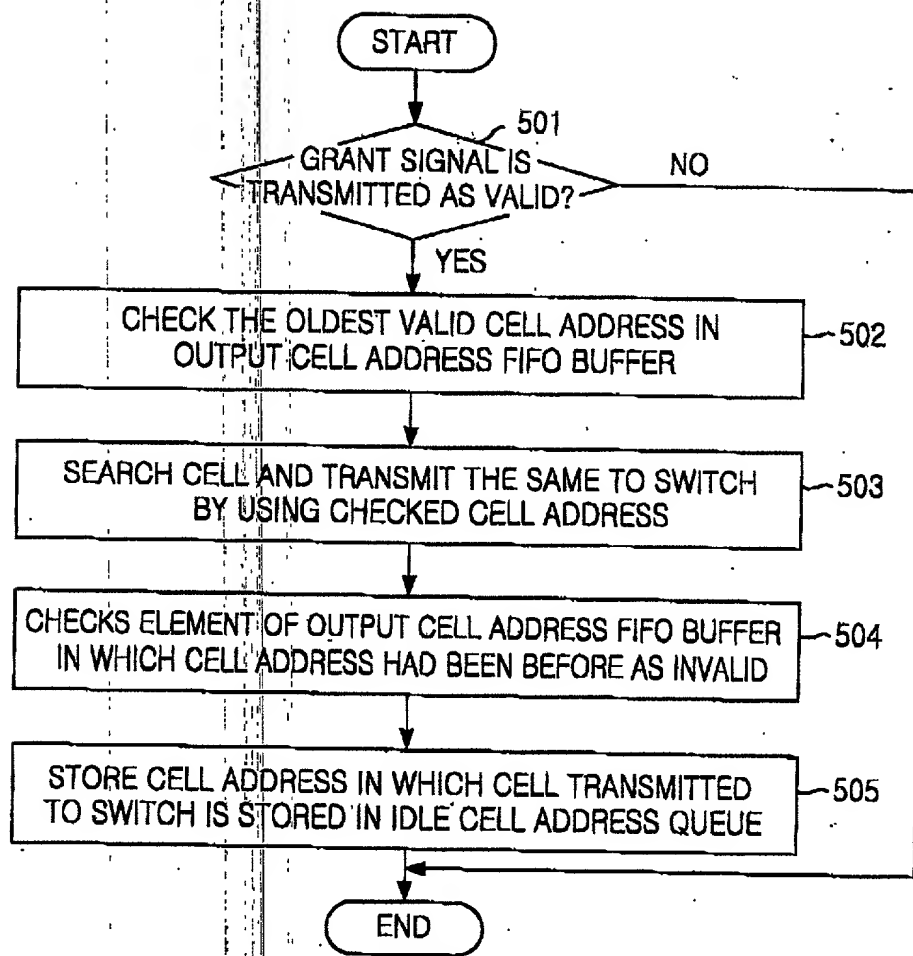


FIG. 6

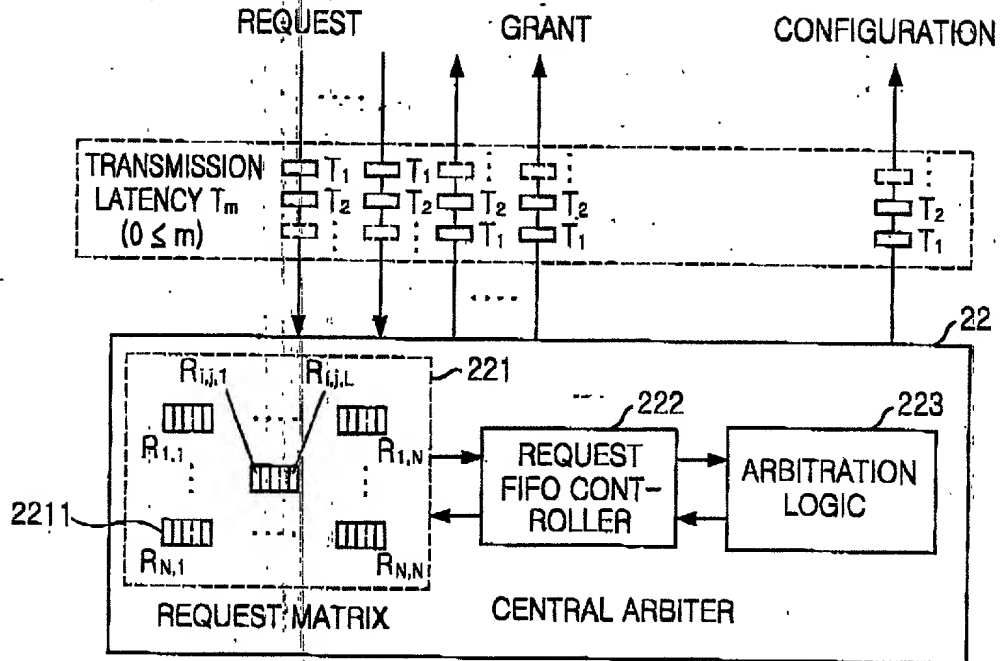
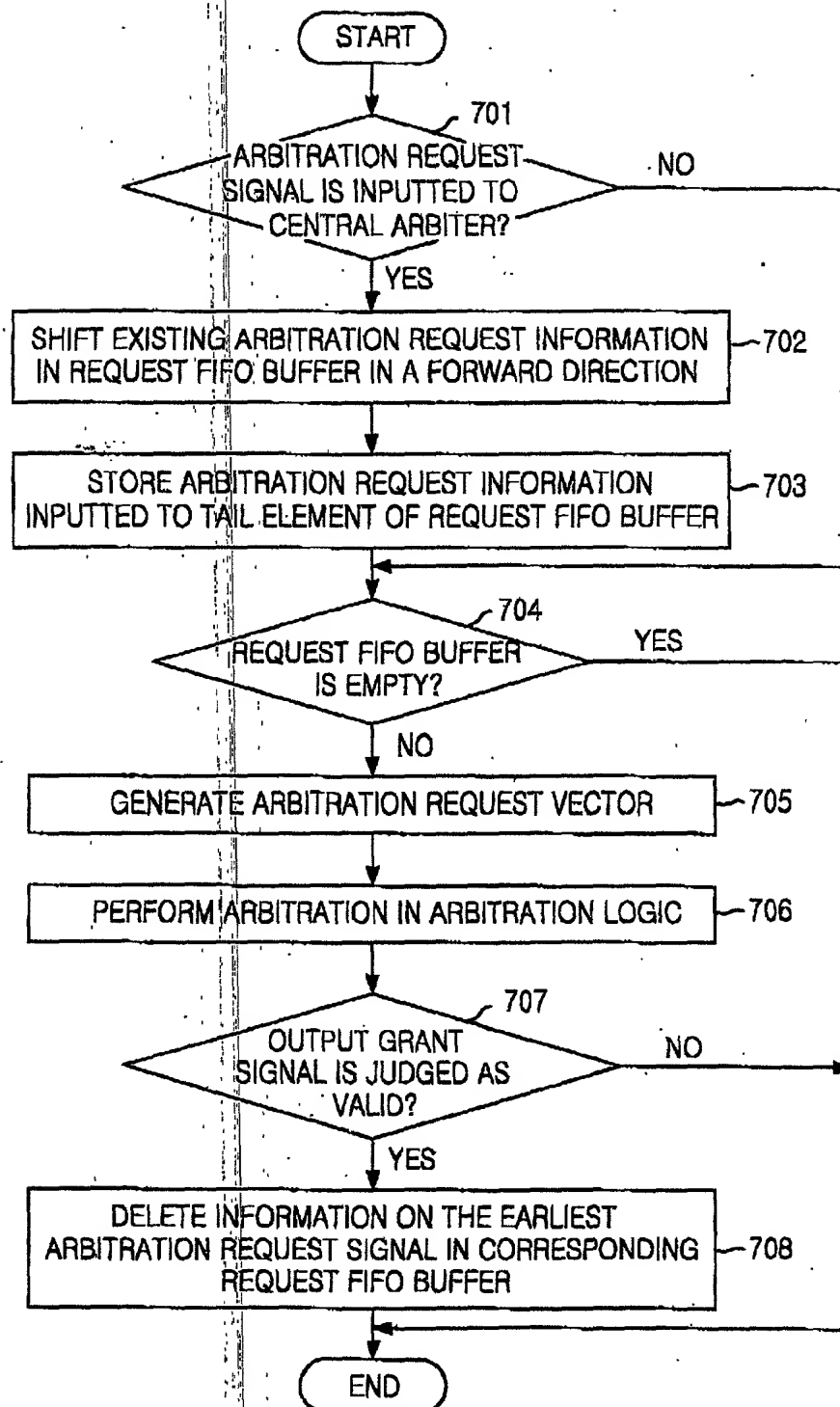


FIG. 7



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FIG. 8

